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Abstract:	<p>Sacrificial anodes are intrinsic to the protection of boats and marine structures by preventing the corrosion of metals higher up the galvanic scale through their preferential breakdown. The dissolution of anodes directly inputs component metals into local receiving waters, with variable rates of dissolution evident in coastal and estuarine environments. With recent changes to the Environmental Quality Standard (EQS), the load for zinc in estuaries such as the Hamble, UK, which has a large amount of recreational craft, now exceeds the zinc standard of 7.9µg/l. A survey of boat owners determined corrosion rates and estimated zinc loading at between 6.95-7.11t/yr. The research confirms the variable anode corrosion within the Hamble and highlighted a lack of awareness of anode technology among boat owners. Monitoring and investigation discounted metal structures and subterranean power cables as being responsible for these variations but instead linked accelerated dissolution to marina power supplies and estuarine salinity variations.</p>	
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Abstract

Sacrificial anodes are intrinsic to the protection of boats and marine structures by preventing the corrosion of metals higher up the galvanic scale through their preferential breakdown. The dissolution of anodes directly inputs component metals into local receiving waters, with variable rates of dissolution evident in coastal and estuarine environments. With recent changes to the Environmental Quality Standard (EQS), the load for zinc in estuaries such as the Hamble, UK, which has a large amount of recreational craft, now exceeds the zinc standard of 7.9µg/l. A survey of boat owners determined corrosion rates and estimated zinc loading at between 6.95-7.11t/yr. The research confirms the variable anode corrosion within the Hamble and highlighted a lack of awareness of anode technology among boat owners. Monitoring and investigation discounted metal structures and subterranean power cables as being responsible for these variations but instead linked accelerated dissolution to marina power supplies and estuarine salinity variations.

Keywords: Sacrificial anode, zinc, EQS, boats, Hamble, estuary

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1. Introduction

1
2 Estuaries are an interface between terrestrial ecosystems and oceans; they are a vulnerable and dynamic
3 environment often subject to wide tidal fluctuations and salinity ranges. Estuaries are also often home to large
4 human populations, ports, marinas and industry (Matthiessen *et al.*, 1999; Sámano *et al.*, 2016). As such, they
5 receive inputs of contaminants from a variety of sources, which need to be quantified and carefully managed to
6 ensure compliance with Environmental Quality Standards (EQS) and the health of the ecology present.
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10 The large amount of zinc released annually from anodes has the potential to contribute to estuaries exceeding
11 the zinc EQS (Bird *et al.*, 1996; Boxall *et al.*, 2000). The new EQS for zinc in UK estuaries is 7.9µg/l of total
12 dissolved zinc, which includes a background level of 1.1µg/l and an EQS of 6.8 µg/l (Maycock *et al.*, 2012),
13 significantly reduced from the previous value of 40µg/l for estuarine waters.
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16 Studies have shown that zinc anodes on leisure craft can be a localised source of zinc to estuaries, which are
17 capable of raising zinc levels in the vicinity of marinas and harbours, for example concentrations of 19.9µg/l of
18 total dissolved zinc were observed at Poole Harbour, which was significantly above control sites of 2µg/l and
19 the new EQS of 7.9µg/l. Similar concentrations were observed on the Hamble and Orwell estuaries (Bird *et al.*,
20 1996; Boxall *et al.*, 2000; Matthiessen *et al.*, 1999). Studies on anode use and dissolution rates were also carried
21 out in the Plymouth area by Wood (2014), Cathery, (2014) and Harrison (2015). Wood (2014) and Harrison
22 (2015) found Plymouth marinas to have higher zinc concentrations in water and sediment samples, compared
23 with nearby control sites. The alternative material to zinc anodes in the marine environment is aluminium (Mao
24 *et al.*, 2011). Aluminium anodes are used less frequently than zinc, mainly due to habits of boat owners and zinc
25 being more commonly used in the past.
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29 Sacrificial anodes are used to prevent corrosion on boat hulls; components on vessels such as propellers and
30 other marine structures including wind turbines, pipelines and pontoons. This electrochemical process is known
31 as cathodic protection (Young *et al.*, 1979; Bird *et al.*, 1996; DeGiorgi *et al.*, 1998; Rousseau *et al.*, 2009).
32 Different elements are used for anodes depending on the ambient environmental conditions. Zinc anodes are
33 most commonly used in seawater, aluminium anodes in seawater and brackish water and magnesium anodes in
34 freshwater (MGDuff, 2016). It is recommended that zinc anodes comply with a US Army Standard (adopted
35 worldwide) of 99% zinc with trace amounts of aluminium, cadmium and iron to make an alloy (Wagner *et al.*,
36 1996). Aluminium and magnesium anodes have similar standards. To function correctly the anode metal has to
37 be lower in the galvanic series than the metal it is protecting, metals lower in the series have a lower potential
38 voltage (Rousseau *et al.*, 2009). Aluminium anodes are not however as popular among boat owners as zinc
39 anodes, this is thought to be due to tradition and habits of boat owners. It is estimated that there are
40 approximately 382,000 marine leisure boats in the UK, many of which will replace anodes and antifouling paint
41 annually (Comber *et al.*, 2002; BMF *et al.*, 2013). The Hamble estuary is home to approximately 3,000 leisure
42 craft alone and is reported to be an area of high anode dissolution by boat owners.
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46 For anodes to cathodically protect a boat or structure the anode has to be in direct contact with the item being
47 protected or be connected by a wire (Rousseau *et al.*, 2009; MGDuff, 2016). Anodes protect the cathode by
48 forming an electrochemical cell. The flow of the electrons in this cell results in accelerated corrosion and
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dissolution of the anode and, in turn, a reduction in corrosion of the cathode (Genesca and Juares, 2000; Rousseau *et al.*, 2009). For this process to remain effective the corroded anode has to be replaced with a fresh anode. The interval of this replacement is generally considered a year, but depends on observed dissolution rates and individual boat owners preferences (Rousseau *et al.*, 2009; Harris, 2008).

A number of reasons have been suggested for this varied anode dissolution rate by boat owners including stray electrical currents, salinity variations, pH of the estuary, current flow, the number of boats present, metal bridges and metal items in the estuary.

Anodes, however, are not the only source of zinc to estuaries. Estuaries with large human populations and boat numbers have the potential to exceed EQS standards due to zinc inputs from sources such as sewage effluent (Gardner *et al.*, 2012), antifouling paints (Rees *et al.*, 2014), road runoff (Rule *et al.*, 2006; Comber *et al.*, 2014), coastal and terrestrial landfill sites (O'Shea *et al.*, 2014) and abandoned mines (Beane *et al.*, 2016).

Within estuaries antifouling paints are also a source of zinc and copper to estuaries with some boats using zinc based paints such as zinc pyrithione, zineb, zinc acrylate copolymers or paints containing zinc additives and pigments (Young *et al.*, 1979; Yebra *et al.*, 2004; Rees *et al.*, 2014; Turner *et al.*, 2014). Zinc pigments often give a yellow colour or zinc oxide white (Abel, 1999). Boats are often abandoned in estuaries worldwide and these can continue to leach metals such as copper and zinc from paints and other metal items into estuaries (Rees *et al.*, 2014; Turner and Rees, 2016). Many antifouling paints are copper based which raises copper concentrations in the vicinity of marinas, harbours and sediments surrounding abandoned boats (Comber *et al.*, 2002; Rees *et al.*, 2014; Briant *et al.*, 2013).

This paper examines the perceived factors for a varied anode dissolution rate by boat owners (Fig 1), to determine if this is the case. This study utilises environmental monitoring data, questionnaire surveys and interviews with boat owners to examine their awareness surrounding anode use. Calculations to show zinc loads to the Hamble estuary from average corrosion rates were determined. The data collected from this research will help to quantify the significance of anodes as a source of zinc within the estuary, identify the factors controlling anode dissolution within the Hamble and beyond and to assist in suggesting mechanisms for future management of zinc in estuarine waters.

2. Method

2.1 Survey of boat owners

A self-administered questionnaire survey was designed and distributed to address the following questions; the types of sacrificial anodes being used; rate of anode replacement, awareness and knowledge of boat owners regarding anode use on the Hamble. Boat owners were also asked what antifouling paint they used to determine if they included zinc based paints. The anode survey was piloted by email to four boat owners; minor adjustments were made, before it was sent out. The survey was distributed via emails, emailed newsletters, twitter, and yachting forums. Boat owners within marinas (with electrical hook-up) and those without electrical supply in mid-channel were approached to determine if they observed different anode corrosion rates. Email

1 yielded the greatest response rate, which along with discussions with boat owners and anode companies
2 suggested reasons for varied anode dissolution rate and also areas of accelerated wear on the Hamble estuary.
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4 **2.2 Anode corrosion rate calculation**

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6 An anode corrosion rate (kg/year/vessel) was calculated for the Hamble from the survey data. The corrosion rate
7 was calculated using the weight of each anode new and the percentage of the anode corroded after one year (the
8 recommenced life for an anode). Where anodes were replaced at longer or shorter time intervals then the weight
9 and percentages were calculated and normalised for a year. An average for the Hamble mid channel mooring
10 berths and marina berths were then calculated (Tables 1 and 2). The corrosion rates were then multiplied by
11 three as studies have found vessels average three anodes per vessel (Wood, 2014); this provided a corrosion rate
12 per vessel rather than per anode
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17 **2.3 Water sampling and analysis**

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19 Water samples were collected in the Hamble estuary (mid channel and marinas) at a depth of 1.2 metres from
20 the mouth, upstream, to the last moorings above Bursledon, in June and October 2015 and January 2016 (Fig 1).
21 All equipment used for water sampling including, Ocean Test water sampler, LDPE bottles, filters
22 and filter units were hydrochloric acid washed (10% analytical grade HCl) and washed copiously with ultra-high
23 purity water (UHP, 18.2M Ω -cm). These samples were filtered through pre-acid washed 0.4 μ m polycarbonate
24 filter membrane (Whatman Ltd) and analysed for their total dissolved zinc concentrations using cathodic
25 stripping voltammetry (Pearson *et al.*, 2016). Salinity measurements were also taken at each sampling site and
26 salinity profiling on separate occasions at high and low tide to determine salinity extremes. .
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33 The water samples were analysed using cathodic stripping voltammetry VA Computrace 797 (Metrohm). Total
34 dissolved acidified zinc samples (ca. 30ml, 90 μ l hydrogen peroxide) were prepared for analysis by UV
35 irradiation for 4 hours by a 400W medium pressure Hg lamp with Photochemical Reactors. 10ml aliquots of
36 sample with a 100 μ l addition of a 238.3 μ g/l concentration of 4-(2-hydroxyethyl)-1-piperazineethanesulfonic
37 acid (HEPES) and a 25 μ l addition concentration of 164.3g/l ammonium pyrrolidine dithiocarbamate (APDC)
38 were purged for 3 mins with N₂ to degas the sample. Deposition times of 5-30 seconds were used depending on
39 the zinc concentration present in the sample, at a potential of -0.9V. Stripping was done by employing a sweep
40 potential between -1.2V and 0.9V (Pearson *et al.*, 2016).
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46 **2.4 XRF analyses of new zinc anodes**

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48 A number of new zinc anodes were analysed for their metal content using a portable XRF (Niton XL 3T Gold
49 Plus). The following parameters were used for the XRF, standard all metals, 180 second measurement time. One
50 reading was taken from each anode in the majority of cases. Some of the anodes were from the same batch of
51 anodes and were analysed to determine variations within anode quality.
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3. Results and Discussion

3.1 Exceedance of EQS on the Hamble

The Hamble estuary frequently exceeds the EQS of 7.9µg/l, with total dissolved zinc concentrations ranging between 3-31µg/l, with the majority of sites being between 5 -20µg/l across June and October 2015 and January 2016 (Fig 2). The EQS is based on an average concentration. The average across the three sampling sessions for the Hamble ranges between 4.14- 16.74µg/l. 10 out of the 22 sampling sites fail the EQS based on average values, the majority of which are mid estuary samples (Fig 2). Higher zinc levels occur in the mid estuary due to high boat numbers in this section of the estuary, as well as a wastewater treatment works at Bursledon, anodes are however, likely to be the largest source of zinc to the Hamble estuary due to 3,000 being present (Fig 2). It is thought that most of the zinc from anodes is deposited into sediments, biota or flushed from estuaries (Matthiessen *et al.*, 1999; Bird *et al.*, 1996). There is however still sufficient concentrations within the water body of the estuary to raise levels above the EQS.

Studies in the Plymouth region, Poole Harbour, Orwell and Blackwater (Matthiessen *et al.*, 1999; Boxall 2000; Wood 2014; Cathery, 2014; Harrison, 2015) have shown large amounts of zinc are released into other estuaries each year from anodes, estuaries with large boat numbers like the Hamble are likely to exceed the 7.9µg/l EQS due to increased zinc from anodes and sources such as wastewater treatment works (Maycock *et al.*, 2012; Gardner *et al.*, 2012). Wood (2014) found marinas in the Plymouth region had total dissolved zinc levels between 2-15µg/l and Harrison (2015) found concentrations of 9- 48µg/l in the same area, which also exceed EQS. This was also observed by Bird (1996) and Boxall (2000) where marinas in some cases exceeded the EQS with observed values on the Hamble of 2.97-9.18µg/l, Poole Harbour and the Orwell 2.16-19.9µg/l. Enclosed marinas with limited tidal flushing and therefore greater water residence times allowing a build-up of zinc in the water, were found to have higher zinc concentrations than open and semi enclosed marinas (Bird 1996; Boxall 2000; Wood 2014; Harrison 2015). Marinas on the Hamble tended to be open or semi enclosed, allowing regular and significantly more tidal flushing of the water. The zinc levels in the Hamble tend to be of the same magnitude as those reported in open and semi enclosed marinas within other estuaries. The marinas and mid channel tend to have similar zinc levels due to the large number of boats being moored in the channel and good tidal flushing of the marinas occurring.

Zinc anodes are not however the only source of zinc to estuaries such as the Hamble with contributions from the riverine section of the catchment, wastewater treatment works effluent and other sources such as road runoff and antifouling paints contributing to EQS levels not being met (Yebra *et al.*, 2004; Rule *et al.*, 2006; Rees *et al.*, 2014; Gardner *et al.*, 2012). Monitoring data, however, shows that riverine concentrations of total dissolved zinc are lower than those measured in the estuary (Fig 2), with 3.24 and 2.83µg/l from two freshwater tributaries of the Hamble on Curbridge Stream. Concentrations of 4.68 µg/l at Curbridge and 1.51µg/l at Botley were measured at the tidal limits of the Hamble's two tidal tributaries with a salinity of 5.44. This is considerably lower than values of 20µg/l at Bursledon with salinities between 25 and 32. The upper estuary is compliant with the EQS. The high salinities around Bursledon show the riverine input is limited. There are two sewage works on the tidal Hamble at Bursledon and downstream of Botley, other rural properties have their own septic tanks.

The population of the Hamble is, however, low so inputs of zinc from sewage effluent discharged to the estuary is considered an insignificant source compared with anode inputs.

3.2 Zinc anode corrosion rates and overall loads to the aquatic environment

Using the calculated corrosion rates (Table 1) and average number of anodes per vessel (3) from the survey and the number of berths within each marina and mid channel moorings it was possible to calculate a total input of zinc to the aquatic environment from anodes (table 2). Boat owners suggested that marinas had a higher anode decay rate than mid channel moorings, hence a corrosion rate calculated for both. The mid channel and marina corrosion rate totals were added together to give an overall Hamble zinc load.

Table 1 Zinc anode release rate per vessel

Release rate location	Berths	Corrosion rate (kg/yr/vessel) (3 anodes)	Load to estuary (t/yr)
Hamble Mid channel (MC)	600	1.62	0.97
Hamble Marina	2400	2.49	5.98
Hamble (Boxall <i>et al.</i> , 2000)	3000	2.37	7.11

The zinc load calculated for mid channel moorings was 0.97t/yr and marinas 5.98t/yr which in total was 6.95t/yr (Table 2). Boxall *et al.*, (2000) also calculated a zinc release rate for the Hamble of 2.37kg/yr for 3 anodes which gave a zinc load of 7.11t/yr. This was calculated for the Hamble as a whole and not mid channel and marinas moorings being separate. This study and Boxall *et al.*, (2000) gave similar zinc corrosion rates' for the Hamble. This shows the high levels of zinc released into the Hamble estuary each year from anodes, which indicates why anodes are the biggest source of zinc to the Hamble. This leads to the EQS not being met. Inputs from individual marinas ranged between 0.12t/yr at RK marine to 0.90t/yr at Mercury Marina (Table 2). Individual vessels in marinas and moorings accounted for between 0.02 -0.90t/yr of zinc to the Hamble.

Table 2 Hamble zinc load calculations from anode release rates

	Corrosion rate (kg/yr/vessel)	1.62	2.49
Hamble estuary moorings	Berths	Zinc load (t/yr)	
Mid Channel	600	0.97	
Hamble Point Marina	230	0.57	
Harbour Board Mooring	10	0.02	
Royal Southern Yacht Club	110	0.27	
Port Hamble	340	0.85	
Mercury Marina	360	0.90	
Universal Marina	250	0.62	
Stone Pier Yard	67	0.17	
Elephant Boatyard	70	0.17	
Swanwick Marina	330	0.82	
Deacons Marina	160	0.40	
Cabin boatyard	73	0.18	
Riverside boatyard	100	0.25	
RK Marine	50	0.12	
Foulkes Boatyard	180	0.45	
Eastland's Boatyard	70	0.12	

Total	0.97	5.98
Hamble Total	6.95	

Corrosion rates of 1.2, 1.35 and 0.81kg/yr/vessel were calculated by studies in the Plymouth area by Wood, (2014), Cathery, (2014) and Harrison (2015), these gave zinc loads ranging between 1.33 – 2.33t/yr for the 1654 vessels in the Plymouth marinas. Corrosion rates from Plymouth marinas were either calculated through a survey (Wood, 2014) or estimates from the MAM-PEC model (Cathery, 2014; Harrison, 2015). This was lower than the Hamble due to lower release rates and boat numbers in Plymouth marinas.

The Orwell and Blackwater estuaries have 1919 and 1860 vessels respectively making them most similar in vessel numbers to the Hamble and Plymouth. Using Corrosion rates of 0.87kg/yr and 2.37kg/yr from Matthiessen *et al.*, (1999) and Boxall *et al.*, (2000) respectively and for an average of three anode per vessel, anodes release between 1.66-4.54t/yr on the Orwell estuary and 1.61t/yr for the Blackwater estuary. Total dissolved zinc levels at Fullbridge on the Blackwater an area of high boat density had an average of 15.23µg/l in 2015 (Environment Agency, 2016). Total dissolved zinc on the Orwell downstream of Wolverstone marina and Pin Mill was 9.51µg/l (Environment Agency, 2016). The Plymouth area and the Blackwater and Orwell estuaries like the Hamble show the large amount of zinc released annually from anodes, which can contribute to EQS levels not being met in estuaries generally.

Matthiessen *et al.*, (1999) used the zinc corrosion rate equation (equation 1) to estimate zinc loads from anodes into Suffolk and Essex estuaries (Stour, Orwell, Crouch, Roach and Blackwater), with the assumption of one anode present per vessel instead of an average of three used in this study. Individual docks accounted for an estimated zinc range of 60 – 540 kg/yr and marinas and moorings a range of 0.29- 320kg/yr. Felixstowe Docks which is a large container port was the largest source of zinc in the area (Matthiessen *et al.*, 1999).

3.3 Questionnaire results and anode replacement rates

Seventy three survey responses were received from the Hamble boat owners, mid channel mooring holders accounted for 58 responses and marinas 15. The survey indicated a number of issues such as the lack of awareness among some boat owners about the use of anodes. In some cases anodes containing inappropriate metals were used; with magnesium used in seawater in two cases on the Hamble. The magnesium will effervesce and quickly corrode in a seawater environment, which has the potential to cause detrimental corrosion to the boat (MGDuff, 2016).

Questionnaire responses show that 59% of boat owners replace their anodes annually on the Hamble (Fig 3). The overall mean replacement for marinas is slightly higher than the moorings mid channel (1.3 compared with 1.2 years respectively, based on the assumption that replacement is at the midpoint of the survey response – i.e. < 1 year = 6 months, >1 to >2 years = 1.5 years, etc.) (Fig 4). A t-test was conducted to compare frequency of anode replacement between mid channel (M = 1.19, SD = 0.47) and marina moorings (M = 1, 28, SD = 0.89), which indicated no significant difference in replacement frequency (t(72) = 0.40 p > 0.05). Although anodes lasting much beyond a year indicate they may not be bonded correctly or are ineffective at preventing corrosion. In some cases boat owners specified they did not know why they used anodes, this was the case in

1 some of the marinas with long replacement times, which increases the average replacement for marinas, as
2 indicated by the following quotes from boat owners on the Hamble“...mine seem to erode very slowly, I changed
3 them this year more out of guilt than needing to...” and “To be honest, I don't really know what they are for and
4 I don't think they are of any significance for me”. This implies a lack of knowledge and awareness among boat
5 owners regarding anode replacement and dissolution mechanisms. The corrosion rates from this study indicates
6 that zinc anodes corrode quicker in marinas compared with mid channel moorings, with the Hamble mid
7 channel corrosion rate of 1.62kg/yr being lower than the Hamble marinas release rate 2.49kg/yr, this is not,
8 however, supported by the replacement rate (Fig 3). This could be due to the number of marina responses and
9 some of the marina questionnaire replies noting anode replacement at relatively large timescales (longer than
10 deemed the norm by the boating community). Boat owners, however, have noticed marinas generally do corrode
11 quicker than mid channel moorings “There is definitely an increase in degradation when in marinas and when
12 on extended periods of shore power. Much less on the mid river moorings” and “Keeping a boat away from
13 marinas is helpful as there are too many stray electrical currents that can eat away anodes and props”.

14
15 From survey responses on the Hamble, boats had an average of three anodes, mainly attached to the propeller
16 shaft and hull, a few vessels had as many as 10 to 15 anodes, however. An average of three anodes per boat was
17 also found in marinas within the Plymouth area by Wood (2014). There was a general trend of larger vessels
18 having greater anode numbers, but some vessels had as many as 8 - 10 anodes at 6 metres as well as up to 14
19 metres (Fig 5). The boat owners with more anodes on smaller vessels had generally experienced corrosion issues
20 so used more anodes. This once again suggest a lack of knowledge and awareness as more anodes on the same
21 metal item will protect it better than one, placed correctly. High anode corrosion rates suggest a possible fault
22 with boat wiring or the wrong size anode being used on the vessel (Harris, 2008).

23
24 There are few other datasets with which to compare these values for the Hamble, however a similar
25 questionnaire survey was undertaken by Wood (2014) on 40 boat owners from Sutton Harbour and Queens
26 Anne Battery in Plymouth. These are high salinity (typically >30) shore side purpose built marinas, with Sutton
27 Harbour having a lock gate to retain water. Based on boat owner responses anode replacment rates were slightly
28 less frequent than in the Hamble (mean 1.7 years). The Plymouth survey suggested the use of shore power
29 accelerated anode wear, just like it did on the Hamble (Wood, 2014). This suggests that there is not just
30 accelerated wear in the Hamble, but what boat owners on the Hamble deem as acclerated wear within marinas
31 with electrical hook ups is further geogrphically spread than just on the Hamble. Electrical hook ups are not
32 present at mid channel sites so are not an issue.

33
34 Time in the marina or at mooring showed no strong relationship with anode replacment frequency in marina
35 locations on the Hamble. Vessels tend to stay in marinas or at moorings year round and are only coming out of
36 the water for maintenance which often includes anode replacment and having antifouling paint re-applied (Fig
37 6). This maintenace work is normally only a matter of weeks and is generally in the winter months. The
38 number of vessels out of the water at anyone time is minimal compared with the total numbers present on the
39 Hamble.

3.4 Possible factors controlling anode corrosion rates

The survey indicated that knowledge and awareness surrounding anodes among some boat owners is limited. The science behind cathodic protection is complicated, so in a bid to explain and understand it, non-scientific theories have been suggested. Many of these theories appear plausible but chemically are not possible, these theories spread through hearsay rather than fact. The idea that metal objects in the river and the metal railway bridge seem plausible, as anodes would cathodically protect many of these items. However as boats are not physically connected to these items an electrical cell will not be formed, meaning anodes cannot corrode via this source (Rousseau *et al.*, 2009; Wagner *et al.*, 1996). This therefore means metal debris, D-Day shell cartridges, scrap aluminium possible thrown in the river from the Fairey Aircraft factory, mooring debris in the estuary and the railway bridge, all suggested by boat owners as reasons, will not effect and cause variation in anode dissolution rates. They could however increase metal levels within sediments and the water column of the Hamble, which in turn will impact on compliance with the EQS.

Variations in salinity are likely to be one of the main factor controlling anode corrosion rates. Low salinity waters causes a calcareous build up on the anode which effects the rate and ability of corrosion (Rousseau *et al.*, 2009; Caplat *et al.*, 2010). Zinc anodes are created from an alloy (Table 3), which allows a steady rate of corrosion in marine waters, this is a US Military Standard. These standards are created for seawater conditions, which means they are less effective in brackish waters and zinc anodes do not corrode in freshwater, due to a layer of zinc hydroxide ($Zn(OH)_2$) forming on the anode, this will form to a lesser extent in brackish waters (Wagner *et al.*, 1996; Gavrila *et al.*, 2000; Jelmert and Van Leeuwen, 2000; Harris, 2008). Freshwater is 10 times less conductive than seawater, zinc (-0.98 to -1.03V) corrodes at a higher voltage than magnesium (-1.60 to -1.63V) so is better suited to seawater (Morgan, 1987). If zinc anodes are removed from water they coat over with a layer of zinc hydroxide and calcium which prevents corrosion, this can occur if boats moorings dry out at low water (Gavrila *et al.*, 2000) (Fig 7). Most boats which have indicated a varied and accelerated dissolution rate are moored at Bursledon or upstream of Bursledon on the Hamble estuary. The salinity in this area ranges between 17 and 31, but can be as low as 5 at low water in extreme conditions (Environment Agency, 2016) (Data collected during this study). The salinity range that occurs in this area is likely to be a reason for the suggested accelerated wear. Yachting forums and discussions with boat owners indicated areas of accelerated corrosion also occurred beyond the Hamble, with marinas on the in the Solent, river Orwell in Suffolk and Chatham marina in Kent being suggested. Salinity and stray currents have been suggested as reasons like on the Hamble. The Orwell estuary below Ipswich Docks had a salinity range of 23.36 – 32.64 during 2015 (Environment Agency, 2016), which is similar to the Hamble estuary. Salinity on the Orwell reached as high as 34 at Shotley (Environment Agency, 2016).

Stray currents is a term often referred to by boat owners as a reason for a varied anode dissolution rates, stray currents in the marine industry are referred to as the portion of current that flows over a path other than the intended path (ACE Group, 2014). Stray current corrosion is where an outside power source flows into the vessels metal components and out through the water as a ground (Fig 8) (Corrosionpedia, 2015). These stray currents can be Direct Currents (DC) which generally cause corrosion to boat fittings and anodes, and Alternating Current (AC) which introduces high voltage current to area surrounding boats, which can result in death, but does not cause corrosion to vessel, this should not occur due to earthing and isolating of these currents

1 at power supplies (ACE Group, 2014). When boat owners refer to stray currents they are referring to DC stray
2 currents. Many marinas provide electrical hook ups powering lighting, power points and cookers on modern
3 boats, these power supplies are AC currents. There are two main types of DC stray currents, external sources to
4 the vessel from marina power supplies and other vessels using the same power supply and internal sources
5 which could be an incorrectly wired appliance or an electrical accessory, anodes act as an earth for static
6 electricity for these appliances, through the engine negative terminal which in turn is connected to the anode
7 (Newboatbuilders, 2015). The boat with the weaker earthing system is most likely to be affected, this could be
8 due to the quality of the wiring, the earthing style, or type of electrical equipment on board. Many boats are
9 wired by boat owners, or terrestrial electricians instead of marine ones, this can lead to grounding and fusing
10 issues due to boat wiring requiring different techniques to terrestrial electricians. Boats in marinas are connected to
11 power supplies and each other and the pontoon through a common earth wire, which allows these stray currents
12 to effect more than one vessel. This can lead to corrosion on a vessel, if it has no anodes, the anodes are not
13 connected correctly or not all metal items are protected by anodes. If the next boat is leaking a DC current the
14 neighbouring boat can act as the DC ground through the marina power supply if both connected to this power
15 source and the other vessel provides the path of least resistance to the ground (Newboatbuilders, 2015;
16 Corrosionpedia, 2015).

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25 A galvanic isolator can be used to break the circuit between vessels, this acts as a filter, blocking the flow of
26 low voltage galvanic (DC) currents but at the same time maintaining the integrity of the earthing circuit
27 (BoatU.S, 2016). About 50% of respondents had these on their vessels moored in marinas and mid channel.
28 Boat owners with these did not however note lower corrosion rates “*I still get some electrolytic corrosion on*
29 *rear propeller when operating fridge on shore power despite the galvanic isolator*”. Galvanic isolators
30 generally do not prevent corrosion and are not designed to, but they will limit it and stop catastrophic corrosion
31 issues if an electric fault occurs, by breaking the two way flow of currents which are a common source of stray
32 currents. Using power supplies only when on board, will also reduce anode corrosion, as vessels are less liable
33 to stray currents without power supplies. It is common for many vessels to be continually plugged into power
34 supplies to maintain battery power, etc. It has also been suggested that some marinas pontoons are not
35 cathodically protected themselves, protection to the pontoon structure is therefore provided by the anodes
36 present on the vessels through the earth wire of the power supplies. A galvanic isolator will also prevent this as
37 allows only one way current flow.

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46 It is likely that anode corrosion is accelerated by stray currents in marinas beyond the Hamble. The Yachting
47 forums indicate this to be the case in many other marinas such as Chatham in Kent and at Wolverstone Marina
48 on the Orwell estuary. This is shown through the following posts: “*Just wondering if any berth holders at MDL*
49 *Chatham maritime, have experienced any unusual corrosion levels either galvanic or leakage / stray currents*
50 *from other boats. Shore power etc?*” and “*I have heard a few knowledgeable people discussing the electrolysis*
51 *at Chatham. It's a bit of an issue apparently*”.

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56 It has been suggested that subterranean power cables under the Hamble (and other estuaries) cause varying
57 anode dissolution rates, this was suggested by two respondents of the survey on the Hamble. These would be
58 protected and earthed so will not release any stray current, as would short out if this was not the case. The
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electrical supply for the railway bridge which has also been suggested as source, this will be isolated and is not connected to the vessels in anyway. This therefore means these will not cause variations in anode corrosion.

XRF analyses on the surface of new anodes was carried out to determine the metal content and to see if anodes meet the US Military standards for anodes (Wagner et al., 1996; Harris, 2008; BoatU.S, 2016). The tests showed that the metal content of anodes in a single batch showed considerable variation and many of these anodes had more metals in then they were meant to. The US Specifications specify that metals not mentioned in the standard should not exceed 0.1% many of these do, with some anodes having as much as 1.1% (Table 3) (Wagner *et al.*, 1996; BoatU.S, 2016). Zinc levels ranged from 96.4 to 98.3%, so all were slightly under the 99.3% standard. It is also thought that some cheaper manufacturer’s anodes generally do not meet these standards. The majority of anodes had only one XRF replicate, but a couple of anodes did have three and these showed very little variation in metal content (Table 4). The US standard is meant to be the best alloy for saltwater environment (Wagner *et al.*, 1996), it is not, however, clear how much these differences will alter anode dissolution. Alloys have been adapted and changed over time (Morgan, 1987), further tests on anodes with different alloy compositions would be needed to test if the variations in anode composition have an effect on the Hamble.

Table 3 XRF analyse of new anodes to indicate metal content present

Element %	Zn anode standard	Bar anode 1	Bar anode 2	Bar anode 3	Hanging anode 1	Hanging anode 2	Hanging anode 3	Hanging anode 4	Hanging anode 5	Hanging anode 6	Hull anode 1	Hull anode 2
Zn	99.32	98.27	98.03	97.01	97.32	98.24	98.24	96.98	97.41	96.4	97.36	96.74
Si	0.07-0.025	1.31	1.27	2.22	0.75	0.75	0.79	1.07		0.88	1.16	1.15
Al	0.1-0.5		0.64	0.63	1.14	1.32	0.87	1.08	1.01	1.5	0.94	0.97
V						0.03		0.03		0.02	0.06	0.11
Cd	0.07	0.04	0.04	0.04	0.04	0.37	0.03	0.02	0.04	0.04	0.03	0.03
Cr						0.85	0.05	0.05	0.05			0.18
W										1.11		0.83
Cu		0.02										
Fe	0.01	0.01	0.02	0.03	0.03	0.02	0.02	0.05	0.04	0.05		
P								0.12				
Ti												
Zr								0.06				
Pb												
Other/other total	0.01	0.02				0.88	0.05	0.26	0.05	1.13	0.06	1.11

Table 4 XRF analyses reps of a hull anode, to show little variation in most elements within an anode

Element %	Zn anode standard	4kg Hull anode R1	4kg Hull anode R2	4kg Hull anode R3
Zn	99.32	96.6	95.74	96.84
Si	0.07-0.025	1.88	1.04	0.89
Al	0.1-0.5	1.59	2.96	1.23
V			0.02	0.05
Cd	0.07			0.04
Cr		0.04	0.1	0.05
W				
Cu				
Fe	0.01	0.03	0.18	0.10
P				
Ti		0.01		
Zr				
Pb				
Other/other total	0.01	0.05	0.12	0.05

PH has also been suggested as a possible reason, the pH in the Hamble is generally around 8 and does not significantly vary. It is very similar to many other estuaries so is unlikely to be a cause of accelerated corrosion. Fast water flow was also suggested by boat owners on the Hamble. Anodes need some water movement to prevent a calcareous and Zn(OH)₂ build upon them (Rousseau *et al.*, 2009), it is however unlikely fast flowing water will cause greater anode corrosion.

Aluminium could be an alternative for some brackish conditions in the upper reaches of the Hamble, as can be used in brackish and seawater (Harris, 2008; MGDuff, 2016), parts of the Hamble varies from brackish to seawater with changes in river flow, tidal conditions and rainfall. Aluminium is considered less of an environmental concern regarding potential toxicity than zinc in marine waters and currently has no EQS set (Harris, 2008; Mao *et al.*, 2011; Gabelle *et al.*, 2012). This indicates that aluminium could be more suited to these areas, as it is of causes less environmental concern and could reduce anode dissolution rates in these areas. The survey and discussions with boats owners indicated that only a small percentage were, however, using aluminium anodes, with only two respondents using. This is partly due to zinc being more commonly used in the past and zinc being recommended over aluminium by anode manufacturers and suppliers in high to mid salinity regions. If aluminium anodes became more commonly used this could reduce zinc levels to below EQS. Aluminium anodes are relatively widely used on marine structures such as wind farms and larger vessels, so can become more widely used on pleasure craft (Gabelle *et al.*, 2012).

A number of boat owners who completed the survey did not have anodes present on vessels. In some cases this was due to lack of knowledge of anode use. But in a few cases the vessels being wooden they did not need anodes. Some older wooden vessels with very few metal components have never had anodes present, so owners tend not to use.

3.5 Significance of marine paints as a source of zinc

The survey indicated that all vessels were using copper based antifouling paints. The antifouling paints are designed for different boat types, speed of travel and conditions (Yebra *et al.*, 2004). The copper based paints will still however contain zinc as pigments and additives to varying levels (Abel, 1999; Yebra *et al.*, 2004; Rees *et al.*, 2014), which will contribute zinc to the estuarine environment. A study by Turner (2010) of paint fragments in recreational UK boatyards showed that some paint chips comprised up to 15% zinc, and up to 35% copper. Zinc released from antifouling paints on the Hamble could be relatively large with 3000 boats present, with zinc as high as 15% present in some of these paints. This source of zinc is however likely to be small compared to the levels released by zinc anodes each year.

4. Conclusion

This study has shown that a lack of awareness among some boat owners can often lead to false theories about anode dissolution rates and quicker anode dissolution rates. It is likely that stray currents internal and external to the vessel, along with salinity variations and variations in anode quality are the main reasons for a varied anode dissolution rate.

Calculations from the survey indicate that between 6.95-7.11t/yr of zinc are released annually on the Hamble. The Hamble is currently exceeding zinc EQS levels with concentrations as high as 31 µg/l. Anodes are likely to be the largest source of zinc to the Hamble estuary and therefore are likely to be the biggest single contribution to the EQS exceedance.

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Figures

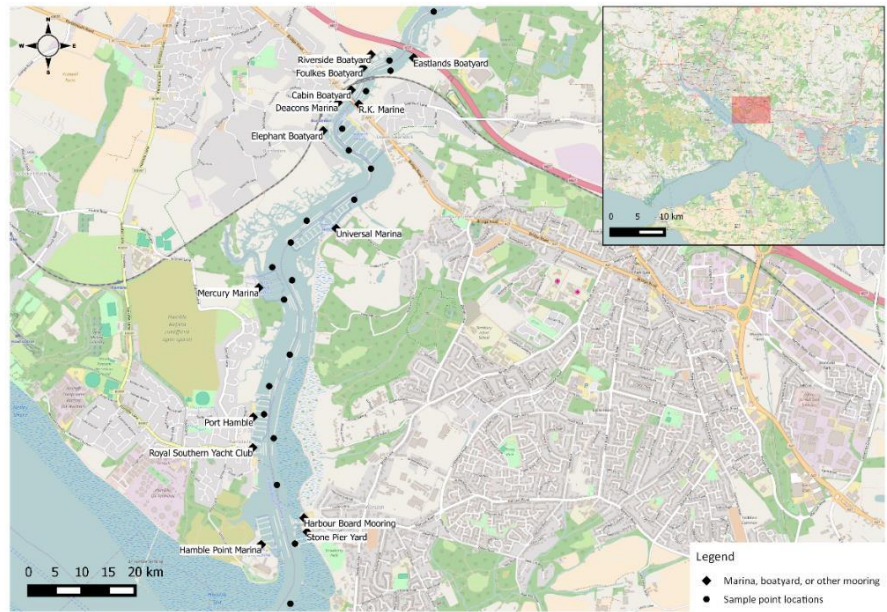


Fig 1 Map showing dissolved zinc sample sites and marinas within the Hamble. The two control sites at Netley (Southampton Water) and Lee on Solent are off the main map

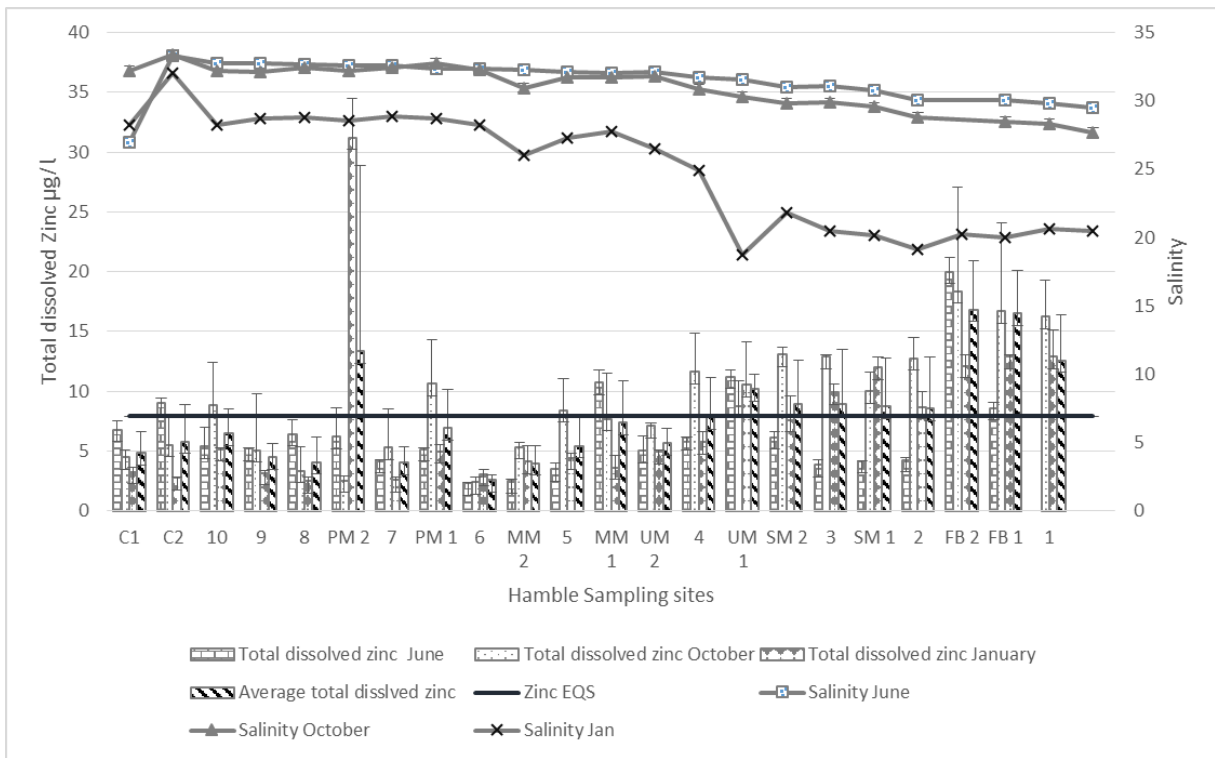


Fig 2 Average total dissolved zinc levels in the Hamble estuary, compared to EQS levels. The salinity plots show the salinity varies from 20-28 in January and 27-32 in June and October. Error bars denote the standard deviation between samples. Site codes 1 – 10 run from the mouth of the Hamble (10) to above Bursledon where boat moorings stop (1), PM is either side of Port Hamble Marina, MM Mercury, UM Universal Marina, SM Swanwick Marina and FB Foulkes Boatyard. C1 control site Netley and C2 control site Lee on Solent, Southampton Water

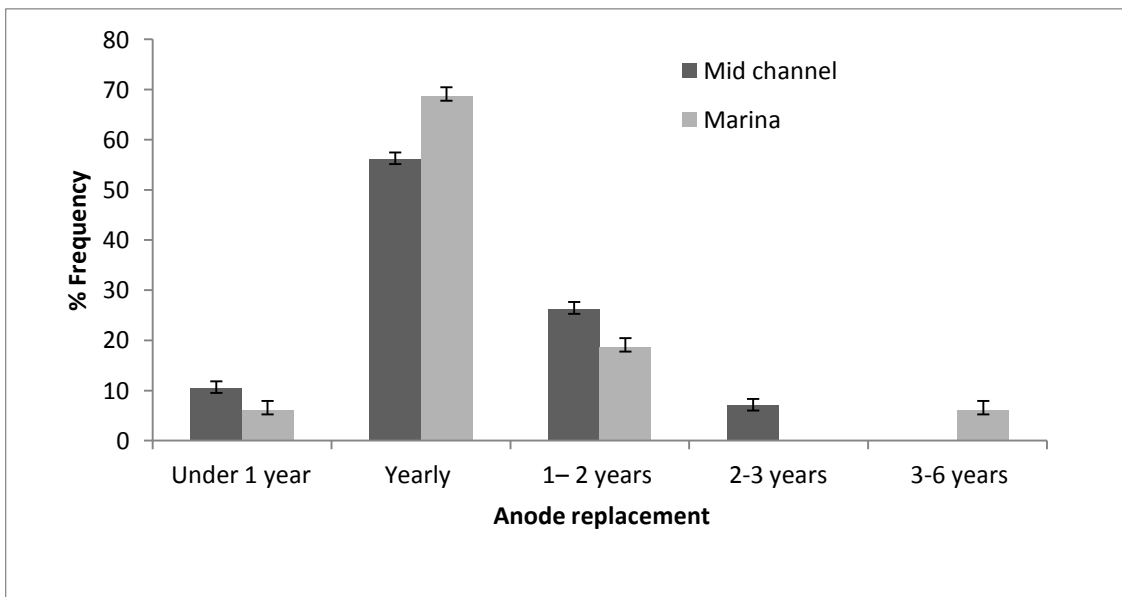


Fig 3 Percentage frequency of anode replacement based on mooring location. Error bars based on 95 percentile values

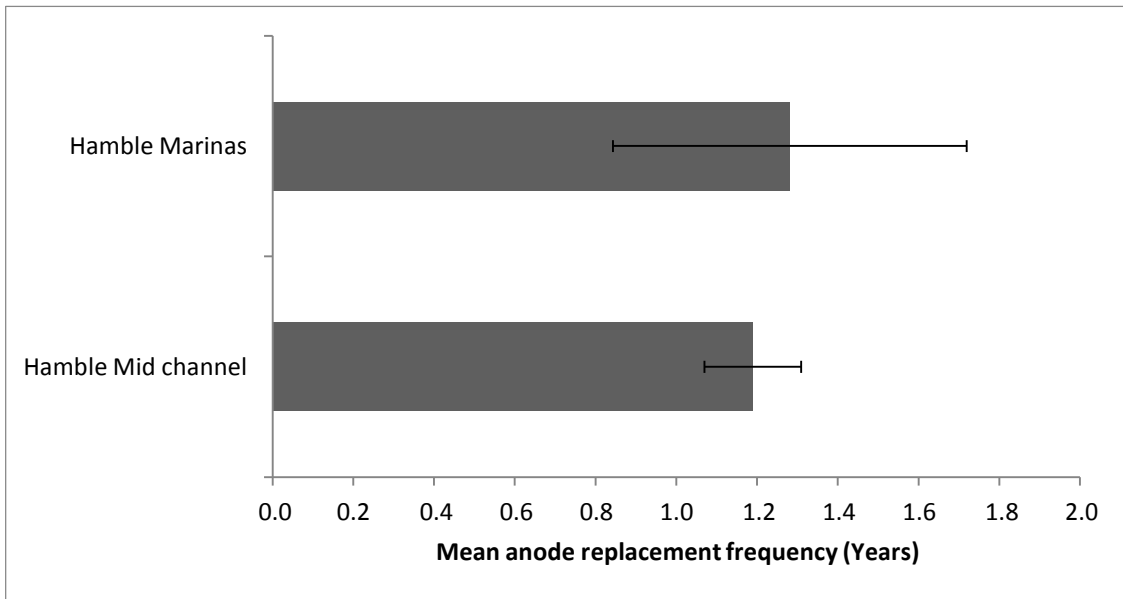


Fig 4 Mean replacement using midpoint frequency replacement in years. Error bars denote 95 percentile confidence intervals

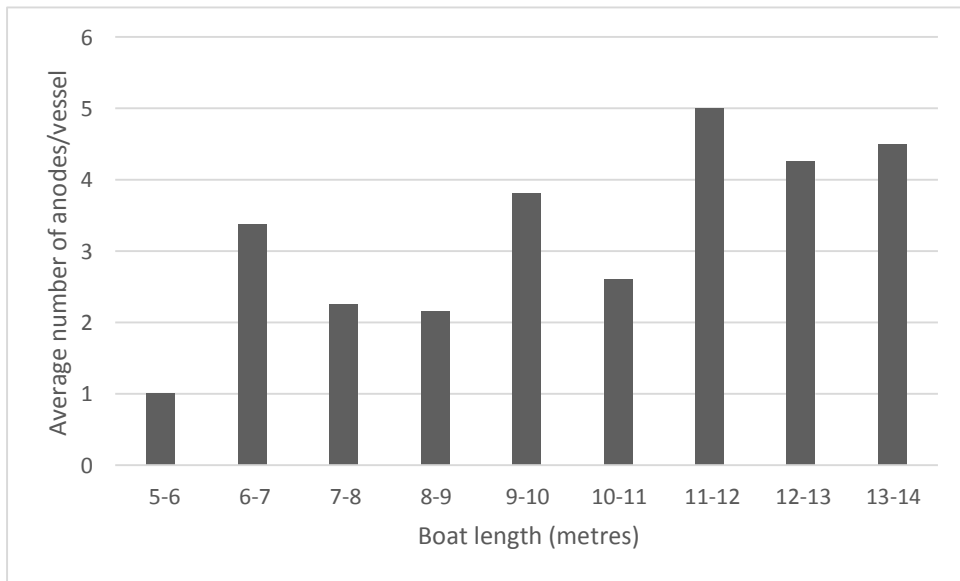


Fig 5 Boat length versus average number of anodes used

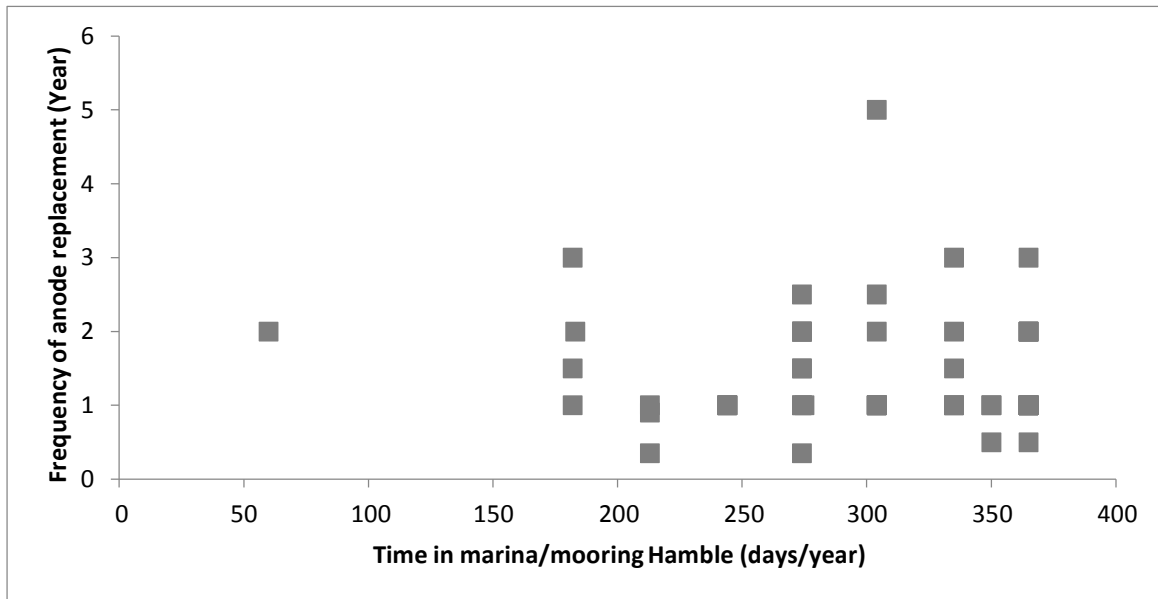


Fig 6 Time in marina versus frequency of anode replacement



Fig 7 Anodes form a layer of zinc hydroxide and calcium if they dry out or are exposed to freshwater or brackish conditions, thus effecting corrosion rates

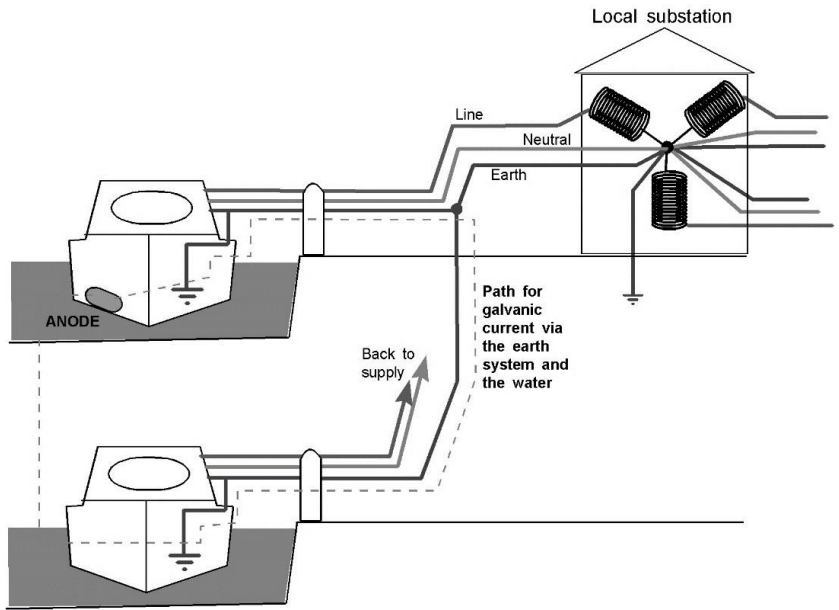


Fig 8 Stray current and galvanic corrosion can occur when vessels are connected through a common earth wire or with a wiring fault on another vessel. This can lead to an accelerated anode dissolution rate (TB-Training, 2012)